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Chapter Test B Chapter: Stoichiometry PART I In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question. 1. Knowing the mole ratio of a reactant and product in a chemical reaction would allow you to determine a. the energy released in the reaction. b. the speed of the reaction.

Name Class Date Assessment) Chapter Test B

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Aspirin production a. 13.5 kg C<sub>9</sub>H<sub>8</sub>O<sub>4</sub> b. 7.66 kg C<sub>4</sub>H<sub>6</sub>O<sub>3</sub> c.

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4.29 L HC<sub>2</sub>H<sub>3</sub>O<sub>2</sub> Ideal stoichiometry calculations do not account for factors that can affect the relative amounts of product chemicals formed in reactions. Input it if you want to ...

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Answer Key Answer:  $4.93 \times 10^5$  L or 49.3 ?L In Example 12.2.1 and Example 12.2.2, the identity of the limiting reactant has been apparent: [Au(CN)<sub>2</sub>]<sup>-</sup>, LaCl<sub>3</sub>, ethanol, and para -nitrophenol. When the limiting reactant is not apparent, we can determine which reactant is limiting by comparing the. Page 5/28.

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Remember it is a MC test, use the answers ... Practice Test Ch3 Stoichiometry (page 2 of 2) 19. The mass of element X found in 1.00 mole of each of four different compounds is 28.0 g, 42.0 g, 56.0 g, and 70 g, respectively. The possible atomic weight of X is a. 8.00 b. 14.0

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